

Chapter 9 Stoichiometry Review Answers

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Q. 4.3 grams of sodium reacts with 2.6 grams of oxygen to produce 5.9 grams of sodium oxide. What's the limiting reactant?

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ANSWER Answer the following questions in the space provided.

1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$
- a. What is the value of the coefficient x in this equation? 40.07 g/mol
- b. What is the molar mass of C_3H_4 ? $2 \text{ mol O}_2 : 1 \text{ mol H}_2\text{O}$
- c. What is the mole ratio of O_2 to H_2O

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Modern Chemistry 77 Stoichiometry CHAPTER 9 REVIEW
Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. _____
The actual yield of a reaction is 22 g and the theoretical yield is 25 g . Calculate the percentage yield.

2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2

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CHAPTER 9 REVIEW. Stoichiometry. MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided.

1. Given the following equation: $C_3H_4(g) + x.O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

a. What is the value of the coefficient x in this equation? b. What is the molar mass of C_3H_4 ? c. How many moles are in an 8.0 g sample of C_3H_4 ? 2. a. What ...

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ANSWER Answer the following questions in the space provided.

1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g)$

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Chemistry Final EXAM Review Chapters 9-16 & Chemistry MATH REVIEW. Chemistry Chapter 9 Test Review Describe a chemical reaction. Define reactant. Define product. Identify the products and reactants in a reaction. Identify a chemical change. Relate the symbols in a chemical equation to the words in a word equation. Write the word equation from a sentence.

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Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 3

PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 ____ The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N₂ are mixed with 120 ...

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Stoichiometry SHORT ANSWER Answer the following questions in the space provided. 1. b The coefficients in a chemical equation represent the (a masses in grams of all reactants and products. (h) relative number of moles of reactants and products. (c) number of atoms of each element in each compound a reaction.

Date. FCHAPJ REV[EW].

Stoichiometry problems can be characterized by two things: (1) the information given in the problem, and (2) the information that is to be solved for, referred to as the unknown. The given and the unknown may both be reactants, both be products, or one may be a reactant while the other is a product.

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CHAPTER 9 REVIEW. Stoichiometry. SECTION 9.2. PROBLEMS

Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:

CHAPTER 9 REVIEW

CHAPTER 9 REVIEW. Stoichiometry. SECTION 1. SHORT ANSWER

Answer the following questions in the space provided. 1. _____
The coefficients in a chemical equation represent the. (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of N₂ are mixed with 120 mol of H₂ according to the following equation: N₂(g) + 3H₂(g)

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Chapter 9 - Stoichiometry Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a...

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