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Chapter 13 (Spontaneous Change: Entropy and Gibbs Energy) is a totally revised version of Chapter 19 from the

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evaluation and analysis tools for mass, energy, entropy, and exergy analysis of open and closed systems.

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Inequality and the Second Law. Entropy and Available Energy. Absolute Temperature. Combined First and Second Laws.

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Standard Entropy,  $S^\circ$ , the absolute entropy of a substance at 1 atm and 25 °C. Units: K mol J.  $\Rightarrow$  Entropy of an element in its most stable form is NOT zero. Trends for Standard Entropy values

- entropy of solid < liquid < gas
- more complex molecules have higher entropies ( $\uparrow$  vibrational energy) E.g.

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## **CHAPTER 18. ENTROPY, FREE ENERGY AND EQUILIBRIUM**

This concept is fundamental to physics and chemistry, and is used in the Second law of thermodynamics, which states that the entropy of a closed system (meaning it doesn't exchange matter or energy with its surroundings) may never decrease. This means that

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the "multiplicity", or number of ways a system can be arranged will never decrease, and that the system will naturally tend to higher disorder.

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## Answers

McMurry, Robert C. Fay, Jill Robinson  
Entropy and the Second Law of  
Thermodynamics First Law of  
Thermodynamics: In any process,  
spontaneous or nonspontaneous, the  
total energy of a system and its  
surroundings is constant.

## **Chapter 17 Thermodynamics:**

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## **Entropy, Free Energy, and ...**

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## **Chapter 18. The Micro/Macro Connection**

Concept 6.1: An organism's metabolism

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transforms matter and energy, subject to the laws of thermodynamics  
Metabolism is the totality of an organism's chemical reactions  
Metabolism is an emergent property of life that

## **Chapter 6 Energy and Life - JU Medicine**

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Chapter 18 Reaction Rates and  
Equilibrium 463 Section Review  
Objectives • Identify two characteristics  
of spontaneous reactions • Describe the  
role of entropy in chemical reactions •  
Identify two factors that determine the  
spontaneity of a reaction • Define Gibbs  
free-energy change Vocabulary Key  
Equation •  $G = H - TS$  Part A Completion Use

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this completion exercise to check your ...

## **05 CTR ch18 7/12/04 8:16 AM Page 463 ENTROPY AND FREE ...**

© 2015 Pearson Education, Inc. Entropy,  
Continued • The net entropy in the  
universe is continually increasing  
(continually running "downhill"). • We  
say net because there are some regions

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in which energy is actually being organized and concentrated. – This occurs in living organisms, which survive by concentrating and organizing energy from food

## **Chapter 18: Thermodynamics**

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The combination of energy decrease and  
entropy increase dictates that

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combustion reactions are spontaneous reactions. Figure 20.2 . Bonfire. A nonspontaneous reaction is a reaction that does not favor the formation of products at the given set of conditions. In order for a reaction to be nonspontaneous, it must be endothermic, accompanied by a ...

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## **Entropy and Free Energy | Chemistry for Non-Majors**

The Second Law of Thermodynamics.

During every energy transfer or transformation, some energy is unusable, and is often lost as heat

According to the second law of thermodynamics. Every energy transfer or transformation increases the

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entropy(disorder) of the universe. ©  
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### **CAMPBELL BIOLOGY**

Chemical Thermodynamics First Law of  
Thermodynamics • You will recall from  
Chapter 5 that energy cannot be created  
or destroyed. • Therefore, the total  
energy of the

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